Good Morning

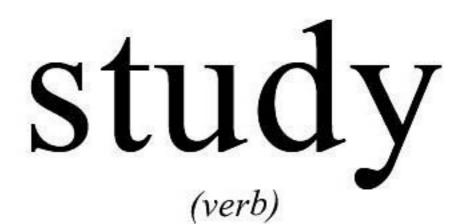
- Please get a whiteboard.
- Write down the type of problem that you would like to see reviewed the most.
- We will begin the review in a moment.

This Week

- M: Review for quiz
- T: Quiz
- W:Thermochemical Equations
- R: Core 3 Thermochemistry
- F: Short Quiz, begin acids and bases

Tonight

Study forQuiz



The act of texting, eating and watching TV with an open textbook nearby.

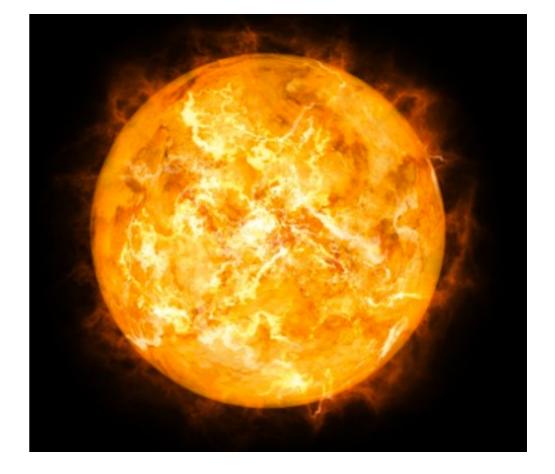
Quiz Topics

- Solutions: M=mol/V & MIVI=M2V2
- Gas pressure and temperature conversions
- Thermochemistry: Heat transfer, specific heat, heat, heating curves, conservation of energy, heat of vaporization, heat of fusion.

A barometer reads ____mmHg. Convert this pressure to kPa. Iatm=14.7psi=760mmHg=760torr=101.3kPa The pressure in a bicycle is ____psi. What is the pressure in the tire in atmospheres?

The temperature on the surface of the sun is 5,778 degrees kelvin. What is the

temperature in celsius?

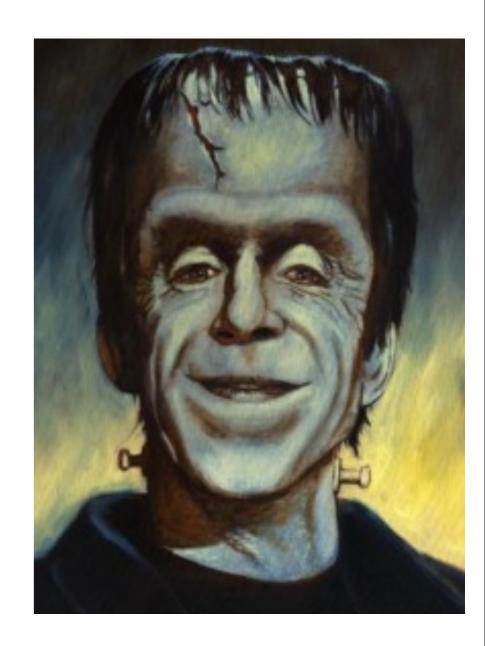


A newly discovered planet has a surface temperature of _____ degrees celsius. What is the temp in kelvin?

Goofy needs to make a ____ml of ___M H2SO4 solution. What mass of H2SO4 does Betty have to measure out?



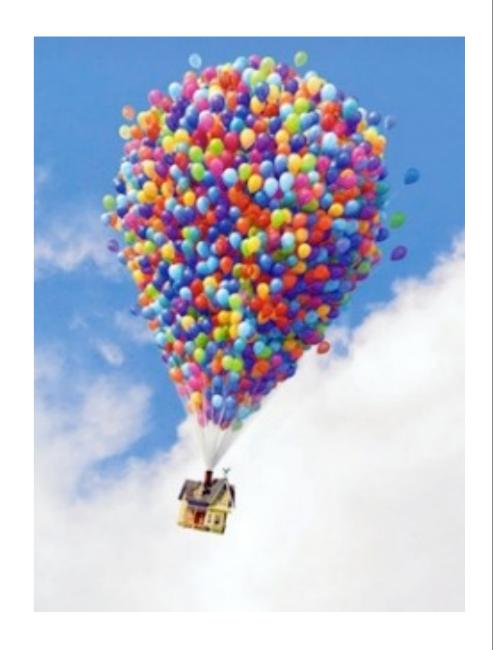
Herman Has ____ml of ___M of ___stock solution. How much water should he add to get ____ml of ___M solution?



Boba Fett creates ____L of ___M HCl to blind security of an escaped con. How many ml of M stock solution did he use?



The specific heat of He is 5.19J/gC. How much energy is released when ____g of He has a temperature change of ____?



An unknown metal has a specific heat of ____J/gC. It absorbs ____J of heat energy. The initial temperature of the metal is ____. What is the final

temperature of the metal?



Calorimetry: Crash Course Chemistry #19

Calculate the heat given off by the system

___ml of water is at ____°C. ___g of unobtainium at ____°C is added to the water. The final temperature of the system is ____°C. What is the specific heat of unobtainium?



The specific heat of granite is 0.79J/g°C. ___g of granite at ___°C is added to ___ml of water. If the water temperature increases by ____°C, what mass of granite was added to the water?

Sulfur melts at 119°C and boils at 445°C. It has a specific heat of sulfur is 0.732J/g°C. ____g of sulfur increases in temperature from ____°C to ____°C. The heat of fusion is 1.73kJ/mol. THe heat of vaporization is 9.8kJ/mol. How much energy is absorbed?